ختبر كيمياء عامه و

gation

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pt

05:34

Question 2

Answer saved

Marked out of
2.00

Flag question

The density of a substance is 1.63 grams per milliliter. What is the mass of 0.25 liters of the substance in grams?

Select one:

O a. 407.5 grams

O b. non of these is correct

● c. 4.075 grams

O d. 68 kilograms

O e. 40.75 grams

Clear my choice

Previous page

N

ختبر كيمياء عامه و

gation

| 3 | 4 | 5 |

mpt

2:05:38

Question 3

Answer saved

Marked out of
2.00

Flag question

Physical properties are used for:

Select one:

O a. a + b + c are correct

O b. Identification of chemical substance

O c. Identification of chemical substances and properties

O d. Identification of chemical properties

● e. All are not correct

Clear my choice

Previous page

N

مختبر كيمياء عامـ

Question 4

Answer saved

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2.00

Flag question

the boiling point has:

Select one:

O a. it is a characteristic of magnitude of the intermolecular force acting between molecules

O b. the entire volume of the liquid has been converted to a gas

O c. all are correct

O d. a + b

● e. temperature of the escaping vapor equal that of liquid

Clear my choice

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Next

مختبر كيمياء ع

Question 5

Not yet
answered

Marked out of
2.00

Flag question

solubility is:

Select one:

● a. the maximum mass (g) of solute dissolve in a non-fixed amount (100 g) of solvent at constant temperature

O b. the maximum mass (g) of solute dissolve in a fixed amount (100 g) of solvent at non constant temperature

O c. the maximum mass (g) of solute dissolve in a fixed amount (100 g) of solvent at constant temperature

O d. all are not correct

O e. the maximum mass (mg) of solute dissolve in a fixed amount (100 g) of solvent at constant temperature

Clear my choice

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search

محبر تيمياء عامه وحصويه

Question 1

Not yet answered

Marked out of 2.00

⚑ Flag question

chemical properties of a substance is:

Select one:

- ⦿ a. depends on its chemical environment
- ◯ b. do not depend on its chemical environment
- ◯ c. depends on its physical environment
- ◯ d. depends on its intermolecular interactions
- ◯ e. all are correct

Clear my choice

Question **3**

Answer saved

Marked out of 2.00

⚑ Flag question

the boiling point has:

Select one:

- ○ a. temperature of the escaping vapor equal that of liquid
- ○ b. it is a characteristic of magnitude of the intermolecular force acting between molecules
- ⦿ c. a + b
- ○ d. the entire volume of the liquid has been converted to a gas
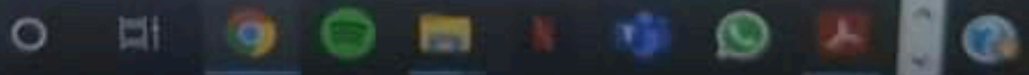- ○ e. all are correct

Clear my choice

Previous page   Next page

the apparatus for determination of boing point has:

Select one:

○ a.

all are correct

○ b.

thermometer

○ c. test tube

○ d. beakers as water bath

◉ e. inverted capillary tube

Clear my choice

مختبر كيمياء عامه

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| 4 | 5 |

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Question 1

Answer saved

Marked out of
2.00

⚑ Flag question
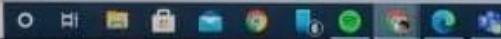
the apparatus for determination of boing point has:

Select one:

○ a. inverted capillary tube

● b.
all are correct

○ c. test tube

○ d. beakers as water bath

○ e.
thermometer

Clear my choice

Next pa

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Question 5

Not yet answered

Marked out of 2.00

⚑ Flag question

NaOH + HCl --> can produce

Select one:

○ a. water or salt

○ b. salt

● c. water and salt

○ d. water and salt and others

○ e. no one is correct

Clear my choice

Previous page   Finish attempt ...

Question **3**

Answer saved

Marked out of 2.00

⚑ Flag question

the boiling point has:

Select one:

○ a. temperature of the escaping vapor equal that of liquid

○ b. it is a characteristic of magnitude of the intermolecular force acting between molecules

⦿ c. a + b

○ d. the entire volume of the liquid has been converted to a gas

○ e. all are correct

Clear my choice

Previous page   Next page

4

the apparatus for determination of boing point has:

Select one:

○ a.

all are correct

○ b.

thermometer

○ c. test tube

○ d. beakers as water bath

● e. inverted capillary tube

Clear my choice

ختبر كيمياء عامة

Question 1

Answer saved

Marked out of 2.00

Flag question

the apparatus for determination of boing point has:

Select one:

O a. inverted capillary tube

◉ b.

all are correct

O c. test tube

O d. beakers as water bath

O e.

thermometer

Clear my choice

Next pa

4

Chemical change is an indicator for the chemical reaction; it could be observed through:

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Select one:

O a. gas evolved

O b. precipitate forming

O c. a or b

◉ d. a or a + b

O e. all are not correct

Clear my choice

Next page

ous page

Not yet
answered

Marked out of
1.00

Flag question

which one is soluble in water

Select one:

- ● a. NaCl

- ○ b. Na2CO3

- ○ c. BaCl2

- ○ d. NH4Cl

- ○ e. all are soluble in water

Clear my choice

Previous page

Question

Not yet answered

Marked out of 2.00

⚑ Flag question

Reagent is:

Select one:

- ⦿ a. a solid chemical or solution have known concentrations of solute

- ○ b. a solid chemical or solution have unknown concentrations of solute

- ○ c. a solid solution only that have known concentrations of solute

- ○ d. a solid chemical only that have known concentrations of solute

- ○ e. all are correct

Clear my choice

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Quiz navigation

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Finish attempt ...

Time left 0:05:52

Question 3

Answer saved

Marked out of 1

⚑ Flag question

If equal masses of $BaCl_2.2H_2O$ and Na_2SO_4 were originally mixed in your unknown mixture, which of the following statements is correct?

MM(g/mol.): $BaCl_2.2H_2O = 244.27$; $Na_2SO_4 = 142.05$; $BaSO_4 = 233.33$

Select one:

O a. The limiting reactant is $BaCl_2.2H_2O$

O b. The limiting reactant is Na_2SO_4

O c. The supernatant solution will give a precipitate upon testing with 0.5 M solution of $BaCl_2.2H_2O$

● d. Both answers (b) and (c) are correct

O e. None of the answers

Clear my choice

Question 4

Answer saved

Marked out of 1

⚑ Flag question

What is the mass of Na_2SO_4 in the mixture if the percentage of $BaCl_2.2H_2O$ is 45.0 % and the mass of the mixture used in the experiment is 1.75 g?

Select one:

O None of the answers

● 0.788 g

O 1.09 g

O 0.693 g

Clear my choice

Previous page

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Answer saved

Marked out of 1

⚑ Flag question

Consider the reaction: $H_2{}_{(g)}$ + $N_2{}_{(g)}$ → $NH_3{}_{(g)}$

Mixing of 1.5 moles of each of hydrogen and nitrogen gases gives how many moles of ammonia (NH_3)?

Select one:

● 1.0 mol.

○ 4.0 mol.

○ 3.0 mol.

○ 2.0 mol.

○ 2.7 mol.
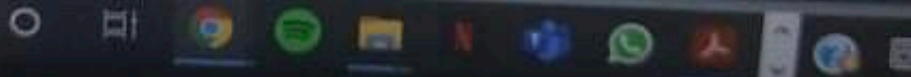
Clear my choice

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which one is soluble in water

Select one:

- ⦿ a. NaCl
- ◯ b. Na2CO3
- ◯ c. BaCl2
- ◯ d. NH4Cl
- ◯ e. all are soluble in water

Clear my choice

Previous page

مختبر

Question 5

Answer saved

Marked out of
2.00

🏴 Flag question

NaOH + HCl --> can produce

Select one:

○ a. water or salt

○ b. salt

◉ c. water and salt

○ d. water and salt and others

○ e. no one is correct

Clear my choice

Previous page

Question 3

Answer saved

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2.5

⚑ Flag question

When 12 moles of O2 reacts with 1.1 moles of C10H8 what is the limiting reactant?

C10H8 + 12 O2 --> 10 CO2 + 4 H2O

Select one:

○ a. C10H8

○ b. Carbon Dioxide

○ c. Oxygen and water

○ d. Water

● e. Oxygen

Clear my choice

Previous page

Nex

Question 1

Answer saved

Marked out of
2.5

⚑ Flag question

2Na3(PO4)2.12.H2O   +   3BaCl2.2H2O   a   Ba3(PO4)2   +   6NaCl + 30 H2O
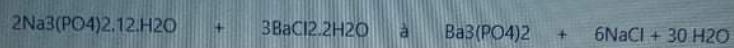
based on this chemical equation of the experiment, Determine the mass of Barium Phosphate that would be predicted using 732.81 g of Barium Chloride Dihydrate and 760.24 g sodium phosphate hydrate

Select one:

○ a. 60.93 g

○ b. 6.093 g

○ c. 0.6093 g

● d. 601.93 mg

○ e. 601.93 g

Clear my choice

What is a limiting reactant?

Select one:

○ a. the reactant that is in excess

● b. all are correct

○ c. the amount of reactants that react with each other

○ d. the product that you can make the most of

○ e. the reactant that determines how much product can be made

Clear my choice

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uiz navigation

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ish attempt...

he left 0:05:27

Question 1

Answer saved

Marked out of 1

⚑ Flag question

Consider the reaction below:

MM(g/mol.): $CH_4=16$, $H_2O=18$

$CH_4 {(g)} + H_2O_{(g)} \rightarrow H_2 {(g)} + CO_{(g)}$

How many grams of hydrogen gas (H_2) are produced when 8 g of CH_4 are mixed with 8 g H_2O?

Select one:

- ○ 2.67 g
- ● 0.89 g
- ○ 3.0 g
- ○ 1.5 g
- ○ None of the answers

Clear my choice

Question 2

Answer saved

Marked out of 1

⚑ Flag question

For the reaction between $BaCl_2.2H_2O$ and Na_2SO_4, to determine the limiting reactant, when the supernatent solution is tested with 0.5 M Na_2SO_4 solution no precipitate was obtained. Accordingly, which of the following statements is correct?

Select one:

- ○ The limiting reactant is $BaCl_2.2H_2O$
- ○ The limiting reactant is Na_2SO_4
- ● None of the reactants is in excess and the mixing between the reactants is stoichiometric
- ○ None of the answers

Clear my choice

Question 2

Answer saved

Marked out of
2.00

⚐ Flag question

NaOH is considered as

Select one:
- ● a. Strong base
- ○ b. Strong acid
- ○ c. Produce H+
- ○ d. Do not react with HCl
- ○ e. All are non-correct

Clear my choice

Previous page

Question 5

Not yet
answered

Marked out of
2.00

⚐ Flag question

NaOH + HCl --> can produce

Select one:
- ○ a. water or salt
- ○ b. salt
- ○ c. water and salt
- ○ d. water and salt and others
- ● e. no one is correct

Clear my choice

Previous page

Finis

Acid material can produce:

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Select one:

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- ● a. H+ in aqueous solution
- ○ b. OH- in aqueous solution
- ○ c. Salt in aqueous solution
- ○ d. Salt and water in aqueous solution
- ○ e. Salt or water in aqueous solution

  Clear my choice

s page

محتوى

The reaction KHP + NaOH à H2O + NaKP is:

Select one:

O a. Mole of KHP equal the mole of NaOH

O b. Mole H+ = mole of OH-

O c. Mole K+ = mole KP-

O d. Mole KHP = mole NaKP

● e. All are correct

Clear my choice

NaOH is considered as

f

Select one:

on
- a. Strong base
- b. Strong acid
- c. Produce H+
- d. Do not react with HCl
- e. All are non-correct

Clear my choice

age

مختبر كيمياء عامه وعضوية

Question **1**

Not yet answered

Marked out of 2.00

⚑ Flag question

The primary standard is

Select one:

○ a. Known high degree of purity

○ b. Large molar mass

○ c. Non-hygroscopic

○ d. React in predictable way

◉ e. All are correct

Clear my choice

Next page

Not yet answered

Marked out of 2.00

⚑ Flag question

Phenolphthalein indicator has

Select one:

○ a. Red − brown color in acid medium

● b. Colorless in acidic medium

○ c. Colorless in basic medium

○ d. Blue color in acidic medium

○ e. All are correct

Clear my choice

Previous page   Next page

Quiz navigation

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Finish attempt ...

Time left **0:07:27**

Question **2**

Not yet answered

Marked out of 2.00

⚑ Flag question

NaOH is considered as

Select one:

- ⦿ a. Strong base
- ○ b. Strong acid
- ○ c. Produce H+
- ○ d. Do not react with HCl
- ○ e. All are non-correct

Clear my choice

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| 1	2	3	4	5

Finish attempt ...

Time left **0:09:03**

Question **5**

Not yet answered

Marked out of 2.00

⚑ Flag question

Acid material can produce:

Select one:

- ⦿ a. H+ in aqueous solution
- ○ b. OH- in aqueous solution
- ○ c. Salt in aqueous solution
- ○ d. Salt and water in aqueous solution
- ○ e. Salt or water in aqueous solution

  Clear my choice

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Finish attempt ...

Time left **0:07:00**